

Introduction to Matter ▪ Section Summary

Describing Matter**Key Concepts**

- What kinds of properties are used to describe matter?
- What are elements, and how do they relate to compounds?
- What are the properties of a mixture?

Matter is anything that has mass and takes up space. **Chemistry** is the study of the properties of matter and how matter changes. In chemistry, a **substance** is a single kind of matter that is pure.

Every form of matter has two kinds of properties—**physical properties** and **chemical properties**. A **physical property** is observed without changing a substance into another substance. Examples of physical properties are hardness, texture, color, and ability to dissolve in water. A **chemical property** is the ability of a substance to change into different substances. Some chemical properties are burning and rusting.

All matter is made up of elements. An **element** is a pure substance that cannot be broken down into any other substance. **Elements are the simplest substances**. Each element is identified by its specific physical and chemical properties. An **atom** is the basic particle that makes up an element. Atoms of most elements can combine with other atoms. A **chemical bond** is the force that holds two atoms together. Atoms often combine to form **molecules**, which are larger particles made of two or more atoms held together by chemical bonds.

When elements are chemically combined, they form compounds having properties that are different from those of the uncombined elements. A **compound** is a pure substance made of two or more elements chemically combined in a set ratio. A compound may be represented by a **chemical formula**. A chemical formula shows the elements in the compound and the ratio of atoms. For example, the chemical formula for carbon dioxide is CO_2 . In carbon dioxide, there are always two oxygen atoms to every one carbon atom.

Elements and compounds are pure substances, but most of the materials you see every day are not. Instead, they are mixtures. A **mixture** is made of two or more substances that are together in the same place, but are not chemically combined. Mixtures differ from compounds in two ways. **Each substance in a mixture keeps its individual properties. Also, the parts of a mixture are not combined in a set ratio.**

A mixture can be heterogeneous or homogeneous. In a **heterogeneous mixture**, you can see the different parts. The substances in a **homogeneous mixture** are so evenly mixed that you cannot see the different parts. A **solution** is an example of a homogeneous mixture. Air is a solution of nitrogen gas, oxygen gas, plus small amounts of other gases. Unlike compounds, mixtures are easily separated into their components. For example, iron filings can be easily removed from salt with a magnet.

Introduction to Matter ▪ Review and Reinforce

Describing Matter

Understanding Main Ideas

Classify each of the following properties by writing physical or chemical on the line.

- _____ 1. Texture
- _____ 2. Ability to react with other substances
- _____ 3. Ability to conduct heat
- _____ 4. Hardness
- _____ 5. Lack of ability to rust
- _____ 6. Physical state

Answer the following questions on the lines provided.

7. How are elements and compounds related?

8. What are two ways in which mixtures differ from compounds?

Building Vocabulary

Match each term with its definition by writing the letter of the correct definition on the line beside the term in the left column.

- | | |
|----------------------------|--|
| _____ 9. matter | a. the basic particle of an element |
| _____ 10. chemistry | b. a mixture in which you can see the different parts |
| _____ 11. substance | c. anything that has mass and takes up space |
| _____ 12. atom | d. a particle formed when two or more atoms combine |
| _____ 13. chemical bond | e. a single kind of matter that is pure |
| _____ 14. molecule | f. a kind of homogeneous mixture |
| _____ 15. chemical formula | g. the study of the properties of matter |
| _____ 16. heterogeneous | h. a mixture in which different parts cannot be seen |
| _____ 17. homogeneous | i. tells the elements and ratio of atoms in a compound |
| _____ 18. solution | j. the force of attraction between two atoms |